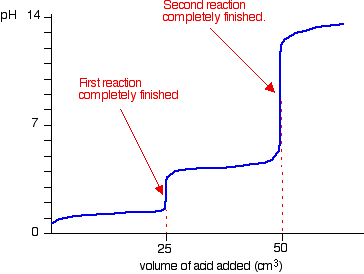
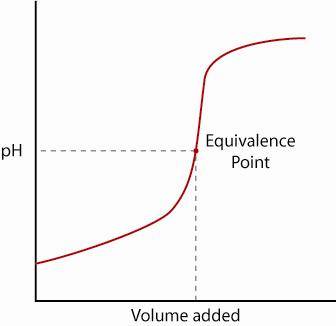
**Acid-Base Review**

*Divide into groups of 3. NO TALKING IS ALLOWED. Each student takes a different coloured pen. One student answers a question and then passes the sheet to the next group member. Group members cannot correct each others’ answers once they have been written. Continue passing the sheet until all questions have been answered.*

1. What is the definition of an Arrhenius Acid? Base?
2. What is the definition of a Bronsted-Lowry acid? Base?
3. Name 2 everyday acids and 2 everyday bases.
4. What is the definition of strong acids and bases?
5. What is the definition of weak acids and bases?
6. List the strong acids.
7. Give an example of a weak acid.
8. List the strong bases.
9. Give an example of a weak base.
10. What is the formula for pH?
11. What is the formula for pOH?
12. What is the pH of HCl with a concentration of 0.1M?
13. What is the pOH of H2SO42- if its concentration is 0.05M?
14. What type of reaction is this: HCl + NaOH 🡪 NaCl + H2O?
15. What is an acid-base indicator?
16. What is the best indicator to use to identify a basic substance around a pH of 8?
17. Is this a diprotic or monoprotic acid?



1. What is the concentration of [H+] and [OH-] ions at the equivalence point shown below?



1. Draw a sketch of the titration of a strong base with a weak acid.
2. What is the pOH if the [H+] is 0.2M?
3. Identify the acid, base, conjugate acid, conjugate base:

HS-(aq) + H2O(l) ↔ H2S(aq) + OH-(aq)

1. What is the [OH-] in a solution with a pH=9?